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Ostwald Process (Manufacture Nitric Acid) Stage 1  $4\text{NH}_3(\text{g}) + 5\text{O}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + 6\text{H}_2\text{O}(\text{g})$  Notes: 1. Temperature =  $900^\circ\text{C}$  2.

Pressure = 4-10 atm 3. Catalyse = Platinum (Pt) Stage 2  $2\text{NO}(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{NO}_2(\text{g})$  Dissolve in water  $3\text{NO}_2(\text{g}) + \text{H}_2\text{O}(\text{l}) \rightarrow$

$2\text{HNO}_3(\text{aq}) + \text{NO}(\text{g})$  Carbon Compound Alkane Alkane

Combustion

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POLITICAL SYSTEM.

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Elements . 1. Reaction with Oxygen . The entire group 1 metal can react with oxygen to form metal oxide.  $4\text{Li} + \text{O}_2 \rightarrow 2\text{Li}_2\text{O}$  .

$4\text{Na} + \text{O}_2 \rightarrow 2\text{Na}_2\text{O}$  .  $4\text{K} + \text{O}_2 \rightarrow 2\text{K}_2\text{O}$  . The metal oxide of group 1 elements can dissolve in water to form alkali (hydroxide) solution .  $\text{Li}_2\text{O} + \text{H}_2\text{O} \rightarrow 2\text{LiOH}$  .  $\text{Na}_2\text{O} + \text{H}_2\text{O} \rightarrow 2\text{NaOH}$  .

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(Average Rate) Rates of reaction = Quantity change of reactants/products Total time for the reaction If the quantity change is immeasurable Rates of reaction = 1 Total time for the reaction Find the Rate From a Graph Average Rate Rates At an Instant The rate of reaction is equal ...

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