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$= 6 \times 1\,000\,000 = 6\,000\,000$ Similarly, the number 1.626×10^{-9} is equivalent to $1.626 \times 10^{-9} = 1.626 \times 0.000\,000\,001 = 0.000\,000\,001\,626$

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$F_2(g) \rightarrow 2 HF(g)$ $\Delta H_f^\circ = -546.6 \text{ kJ mol}^{-1}$

$H_2O(l) \rightarrow H_2(g) + \frac{1}{2} O_2(g)$ $\Delta H_f^\circ = -285.8 \text{ kJ mol}^{-1}$

Calculate The Value

Of ΔH_{rxn}° For $2 H_2(g) + O_2(g) \rightarrow 2 H_2O(l)$

$\Delta H_{rxn}^\circ = -571.6 \text{ kJ mol}^{-1}$

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either in unreacted
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2(aq) produced, and

so [Cl⁻] = 75 mmol / 25

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of $H^+(aq)$ remaining is
equal to the original
number of millimoles
of acid minus the
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